

Video: Kinetic Energy (KE) vs Potential Energy (PE)

**Watch the video, then answer these questions (1-6) in your Interactive Notebook.
You DO NOT need to copy the entire chart, please just answer the question in your notebook.**

For example, for question #1: you will record the definition of Kinetic Energy in your Interactive Notebook.

	Kinetic Energy (KE)	Potential Energy (PE)
Definition	1. (at 0:38)	2. (at 1:45)
Example in cartoon	3	4
Relevance to phases of matter (S, L, G)	Gas particles move fastest, have the highest temperature and Kinetic Energy (KE), and can do the most work.	Potential Energy is stored within covalent and ionic bonds. When the particles of a substance have more space to move (are spread apart), they have more Potential Energy (PE).
When does this energy increase or decrease?	5	6
Example	KE increases if a H ₂ O(l) is heated from 10°C to 30°C, and KE decreases if H ₂ O(l) is cooled from 40°C to 15°C.	*PE increases during phase changes like S-->L and L-->G when energy is added but temperature is not changing. *PE decreases during phase changes like G-->L and L-->S when energy is added but temperature is not changing.

Video: