

Medium Worksheet

Steps:	Draw in the Example for PCl_3
1. Draw out the symbols for the compound. Put the "single" symbol in the center.	
2. Draw the Lewis Dot Diagram for each of the compounds.	
3. Determine how many more electrons each atom needs. Remember that the number of electrons that it needs is the same as the number of bonds that it will add.	
4. Connect the electrons in order to create the bonds.	
5. Check to make sure that each atom has the 8 (or 2 if it is H) electrons.	

Medium Practice - Sketch in your Interactive Notebook

#1: NH_3	#2: H_2O
#3: CO_2	#4: C_2H_4

Challenge Questions - Sketch in your Interactive Notebook!

<p style="text-align: center;">#1: CH_3Cl</p> <p>Hint: C is the central atom with four bonds</p>	<p style="text-align: center;">#2: N_2</p> <p>Hint: a) How many bonds does each atom of N need?</p>
<p style="text-align: center;">#3: CH_2Br_2</p> <p>Hint: C is the central atom with four bonds</p>	<p style="text-align: center;">#4: $\text{CH}_3(\text{OH})$</p> <p>Hints: C is the central atom with four bonds The O atom is bonded to both carbon and the H atom</p>