**Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Medium Neutralization Practice**

**Neutralization in words:**

* An Acid and Base **Always** yields: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Neutralization in chemical formulas:**

* HCl(aq) + NaOH 🡪 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

After **neutralization**, the pH is always closer to**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

Complete and balance each of the acid base neutralizations below.

a. \_\_\_ H2SO4 + \_\_\_ Mg(OH)2 🡪 \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

b. \_\_\_ HI + \_\_\_ KOH 🡪 \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_

c. \_\_\_ H3PO4 + \_\_\_Ca(OH)2 🡪 \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

d. \_\_\_ HBr + \_\_\_ Ba(OH)2 🡪 \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

e. \_\_\_H3PO4 + \_\_\_KOH 🡪 \_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

f. \_\_\_HCl + \_\_\_NaOH 🡪 \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

g. \_\_\_\_ HNO3 + \_\_\_ Al(OH)3 🡪 \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

h. \_\_\_H2SO4 + \_\_\_NaOH 🡪 \_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

i. \_\_\_HClO3 + \_\_\_Al(OH)3 🡪 \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

j. \_\_\_\_HCl + \_\_\_\_Al(OH)3 🡪 \_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_