

Atomic Mass and Total Charge Worksheet Questions

Mild Questions:

1. What is the name of a particle that has a charge of +1? How can you use the name of this subatomic particle to remind you that it has a positive charge?
2. Which subatomic particles are found in the nucleus?
3. Where does the most mass from an atom come from? How do you know?
4. How do you find the total mass of an atom? Why is this the case?
5. Which subatomic particles have a charge?
6. What does it mean to find the total charge of an atom?
7. How much is the mass of an atom if there are 5 protons, 6 neutrons, and 5 electrons? What is the charge? Show your work.

Medium Questions:

1. How do you find the total charge of an atom? Why is this the case?
2. How do you find the total mass of an atom? Why is this the case?
3. What does it mean for an atom to have a neutral charge? If an atom has 4 protons, how many electrons does it need to have a neutral charge?
4. How much is the mass of an atom if there are 5 protons, 6 neutrons, and 5 electrons? What is the charge? Show your work.
5. How much is the mass of an atom if there are 8 protons, 9 neutrons, and 6 electrons? What is the total charge? Show your work.
6. How much is the mass of an atom if there are 7 protons, 6 neutrons, and 5 electrons? What is the total charge? Show your work.

Spicy Questions:

1. How much is the mass of an atom if there are 8 protons, 9 neutrons, and 8 electrons? What is the total charge? Show your work.
2. What happens to the charge of an atom when there are more protons than there are electrons? How do you know?
3. What happens to the charge of an atom when there are more electrons than there are protons? How do you know?
4. An atom has 5 protons. How many neutrons and electrons does an atom have if the mass is 12 and the charge is 0.
5. An atom has 8 protons. How many neutrons and electrons does an atom have if the mass is 15 and the charge is -2.
6. An atom has 4 protons. How many neutrons and electrons does an atom have if the mass is 9 and the charge is +1.